

Download Questions On Net Ionic Equations

PRACTICE PROBLEMS ON NET IONIC EQUATIONS page 1 of 3 Show the complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding! Definitions of molecular, complete ionic, and net ionic equations. All sodium, potassium, ammonium, and nitrate salts are soluble in water. Most chloride salts are soluble except for silver halides. If you aren't sure how to break apart the soluble ionic compounds into the cation(s) and anion(s), you can check out the article on common polyatomic ions. Balance the complete molecular equation. Before writing a net ionic equation, you must first make sure your starting equation is completely balanced. To balance an equation, you add coefficients in front of compounds until there is an equal number of atoms for each element on both sides of the equation. Unit 6 Quiz-- Ionic and Net Ionic Equations: Multiple Choice (Choose the best answer.) ... Which of the answers below is an ionic representation of the reaction: ... What would be the net ionic reaction in problem 7? $K_2 + (aq) + (C_2H_3O_2)_2 \rightarrow 2K^+ + 2C_2H_3O_2^-$ (s)